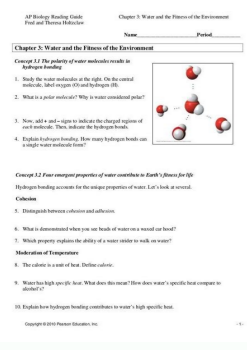


Continue

(KEY)

Water, H₂O			
Electronegativities: H 2.1	O 3.5	Difference 1.4	Polar or nonpolar bond? <u>POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>H-O-H</chem>	<u>BENT</u>	<u>POLAR</u>
Ammonia, NH₃			
Electronegativities: N 3.0	H 2.1	Difference .9	Polar or nonpolar bond? <u>POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>H-N-H</chem>	<u>Trigonal Pyramidal</u>	<u>POLAR</u>
dichloromethane, CCl₂H₂			
Electronegativities: C 2.0	Cl 3.5	Difference 1.5	Polar or nonpolar bond? <u>POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>Cl-C-H</chem>	<u>Tetrahedral</u>	<u>POLAR</u>
Carbon disulfide, CS₂			
Electronegativities: C 2.0	S 2.0	Difference 0	Polar or nonpolar bond? <u>NON POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>S=C=S</chem>	<u>LINEAR</u>	<u>NON POLAR</u>
Boron trichloride, BCl₃			
Electronegativities: B 2.0	Cl 3.5	Difference 1.5	Polar or nonpolar bond? <u>POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>Cl-B-Cl</chem>	<u>Trigonal Planar</u>	<u>NON POLAR</u>
Sulfur dichloride, SCl₂			
Electronegativities: S 2.0	Cl 3.5	Difference 1.5	Polar or nonpolar bond? <u>POLAR</u>
Electron dot diagram	Structural formula	Shape	Polar or Nonpolar Molecule?
	<chem>Cl-S-Cl</chem>	<u>BENT</u>	<u>POLAR</u>



Molecular Formula	Lewis Structure	# of atoms bonded to central atom	# of lone pairs on the central atom	Electron pair Geometry	Molecular Shape	AXE type
CO ₃ ²⁻		3	0	trigonal planar	trigonal planar	AX ₃ E ₀
CO ₂		2	0	linear	linear	AX ₂ E ₀
NH ₃		3	1	tetrahedral	trigonal pyramidal	AX ₃ E ₁
H ₂ O		2	2	tetrahedral	bent	AX ₂ E ₂
PF ₅		5	0	trigonal bipyramidal	trigonal bipyramidal	AX ₅ E ₀
SF ₆		6	0	octahedral	octahedral	AX ₆ E ₀
CF ₄		4	0	tetrahedral	tetrahedral	AX ₄ E ₀
ClF ₃		3	2	trigonal bipyramidal	T shaped	AX ₃ E ₂
XeF ₂		2	3	trigonal bipyramidal	linear	AX ₂ E ₃
SF ₄		4	1	trigonal bipyramidal	seesaw	AX ₄ E ₁
BF ₃		3	0	trigonal planar	trigonal planar	AX ₃ E ₀
ICl ₄		4	2	octahedral	square planar	AX ₄ E ₂

Molecular Geometry and Polarity
CHM 113 Lab Worksheet

Box # _____

Lab Day/Time _____

Student Name(s) _____

Complete Set on Pickup? Yes / No*

Complete Set on Return? Yes / No*

TA Initials: _____

Tetrahedral Electron Domain Geometry (EDG)

**If a set is not complete, collect replacement piece(s)*

**This worksheet does not replace good lab notebook keeping!
ALL INFORMATION RECORDED ON THIS WORKSHEET
SHOULD COME FROM THE BOOK, PRELAB NOTES, THE LAB MANUAL AND
EXPERIMENTAL VALUES/CALCULATIONS YOU HAVE RECORDED IN YOUR
NOTEBOOK!**

These structures all have **4 areas of e⁻ density** and bond angles are approximately 109.5°. Areas of e⁻ density include lone pairs around the central atom and peripheral atoms attached to the central atom. As the number of lone pairs increase around the central atom the bond angles are slightly decreased.

Possible Molecular Geometries from this EDG are: Tetrahedral, Trigonal Pyramidal, and Bent

CF₄

1) Build table to identify the # of bonds:		
Atom	# of VE wanted	# of VE has
C	8	4
F	8	7
F	8	7
F	8	7
F	8	7
TOTAL	40	32
40-32 = 8 shared e ⁻ 8/2 = 4 bonds		

2) Draw the Lewis Structure:

3) How many areas of e⁻ density (lone pairs and atoms) are there?
4

4) What is the Electron Domain Geometry (EDG)? Draw it.

5) What is the Molecular Geometry (MG)? Draw it and indicate bond angles.

What are polar molecules give example. Polarity of bonds and molecules worksheet answers.

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